

Stoichiometry Test Holt

Stoichiometry Test Holt - Holt Chemistry Chapter 9: Stoichiometry Chapter Exam Instructions. Choose your answers to the questions and click 'Next' to see the next set of questions. Holt McDougal Modern Chemistry Chapter 9: Stoichiometry / Practice Exam Exam Instructions: Choose your answers to the questions and click 'Next' to see the next set of questions. Remember it is a MC test, use the answers ... Practice Test Ch 3 Stoichiometry Name _____ Per _____ $2\text{MnO}_2 + 4\text{KOH} + \text{O}_2 + \text{Cl}_2 \rightarrow 2\text{KMnO}_4 + 2\text{KCl} + 2\text{H}_2\text{O}$ 9. For the reaction above, there is 100. g of each reactant available. Which reagent is the limiting reagent? [Molar Masses: MnO Stoichiometry Chemistry Quiz. This is 1 mole of water produced per mole of P_2O_5 . There are 4 moles of water produced per mole of P_2O_5 . There are 6 moles of water produced per mole of P_2O_5 . There are 10 moles of water produced per mole of P_2O_5 . A compound has an empirical formula is NPCl_2 and a molecular weight of 347.66.